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Water quality — Determination of chloride — Silver nitrate titration with chromate indicator (Mohr's method)

*Qualité de l'eau — Dosage des chlorures — Titrage au nitrate d'argent
avec du chromate comme indicateur (Méthode de Mohr)*



Reference number
ISO 9297:1989(E)

Foreword

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Draft International Standards adopted by the technical committees are circulated to the member bodies for approval before their acceptance as International Standards by the ISO Council. They are approved in accordance with ISO procedures requiring at least 75 % approval by the member bodies voting.

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Introduction

Nearly all natural waters, as well as rain water and many waste waters, contain chloride ions. Concentrations can vary widely from a few milligrams per litre in some natural waters to very high concentrations in polluted waste water, marine water and saline ground water.

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1 Scope

1.1 Application range

This International Standard specifies a titration method for the determination of dissolved chloride in water. The method is applicable to the direct determination of dissolved chloride in concentrations between 5 mg/l and 150 mg/l. The working range may be extended to 400 mg/l by using a burette of larger capacity or by sample dilution. Due to many interferences the method is not applicable to heavily polluted waters of low chloride content.

1.2 Interferences

Normal concentrations of common constituents of ground water, surface water and potable water do not interfere with the determination.

The following substances interfere with the method

- Substances forming insoluble silver compounds, such as bromides, iodides, sulfides, cyanides, hexacyanoferrates(II) and hexacyanoferrates(III). If necessary, bromide and iodide ions shall be determined separately, and the result of the chloride determination corrected accordingly.
- Compounds forming complexes with silver ions, such as ammonium and thiosulfate ions.
- Compounds which will reduce chromate ions, including iron(II) and sulfite ions.

The interferences mentioned above will lead to high chloride values.

Highly coloured or turbid solutions may obscure the end point, for example hydrated iron oxide.

Table 1 — Interferences

Substance	Amount interfering mg/l
Br ⁻	3
I ⁻	5
S ²⁻	0,8
CN ⁻	1
Fe(CN) ₆ ⁴⁻	2
Fe(CN) ₆ ³⁻	2
NH ₄ ⁺	100
S ₂ O ₃ ²⁻	200
SO ₃ ²⁻	70
SCN ⁻	3
CrO ₄ ²⁻	1 000
PO ₄ ³⁻	25

Table 1 gives a summary of the concentrations of interfering compounds, in milligrams per litre, that give an increase of approximately 2 % in the result when in the presence of 70 mg/l of chloride.

2 Normative references

The following standards contain provisions which, through reference in this text, constitute provisions of this International Standard. At the time of publication, the editions indicated were valid. All standards are subject to revision, and parties to agreements based on this International Standard are encouraged to investigate the possibility of applying the most recent editions of the standards indicated below. Members of IEC and ISO maintain registers of currently valid International Standards.

ISO 385-1:1984, *Laboratory glassware — Burettes — Part 1: General requirements.*

ISO 5667-1:1980, *Water quality — Sampling — Part 1: Guidance on the design of sampling programmes.*

ISO 5667-2:1982, *Water quality — Sampling — Part 2: Guidance on sampling techniques.*

ISO 5667-3:1985, *Water quality — Sampling — Part 3: Guidance on the preservation and handling of samples.*

ISO 5725:1986, *Precision of test methods — Determination of repeatability and reproducibility for a standard test method by inter-laboratory tests.*

3 Principle

Reaction of chloride with added silver ions to form insoluble silver chloride which precipitates quantitatively. Addition of a small excess of silver ions to form a red brown silver chromate with chromate ions that have been added as an indicator. This reaction is used for indicating the end-point. The pH is maintained in the range of 5 to 9.5 throughout the titration in order to allow precipitation.

4 Reagents

NOTE 1 All silver compounds and solutions are sensitive to light. Silver salts temporarily stain the skin brown.

Use only reagents of recognized analytical grade and only distilled water or water of equivalent purity.

4.1 Silver nitrate, standard volumetric solution, $c(\text{AgNO}_3) \approx 0,02 \text{ mol/l}$.

Dissolve 3,397 4 g of silver nitrate (AgNO_3), previously dried at 105°C in water and dilute to 1 000 ml in a one-mark volumetric flask.

If stored in the dark in a brown glass bottle with glass stoppers, the solution is stable for several months. The solution is standardized against 10 ml sodium chloride standard reference solution (diluted to 100 ml) using the procedure given in 6.1, however, there is no need for pH adjustment.

4.2 Potassium chromate, indicator, 100 g/l solution. Dissolve 10 g of potassium chromate (K_2CrO_4) in water and dilute to 100 ml.

4.3 Sodium chloride, standard reference solution, $c(\text{NaCl}) = 0,02 \text{ mol/l}$. Dissolve 1,168 8 g of sodium chloride (NaCl), previously dried at 105°C , in water and dilute to 1 000 ml in a one-mark volumetric flask.

4.4 Nitric acid, $c(\text{HNO}_3) \approx 0,1 \text{ mol/l}$.

Stored in a glass bottle, the solution is stable indefinitely.

4.5 Sodium hydroxide, solution, $c(\text{NaOH}) \approx 0,1 \text{ mol/l}$.

4.6 Reagent, for improvement of the buffer capacity.

Calcium carbonate (CaCO_3) or sodium hydrogen carbonate (NaHCO_3) in powder form.

5 Apparatus

Ordinary laboratory equipment and

5.1 Burette, of capacity 25 ml, complying with ISO 385-1.

6 Procedure

For sampling and preservation of samples refer to ISO 5667-1, ISO 5667-2 and ISO 5667-3.

6.1 Titration

Pipette a test portion of 100 ml, or a smaller volume of the sample diluted to 100 ml (volume V_a), into either a white porcelain basin, or a conical flask or a beaker held against a white background.

If the pH of the sample is outside the range of 5 to 9.5, adjust the pH using nitric acid (4.4) or sodium hydroxide (4.5) as appropriate, and note the volume required.

If ammonium ions are present in the sample in concentrations exceeding 10 mg/l, adjust the pH to between 6.5 and 7.

Adjust the pH in one aliquot, then take another and, this time without measuring the pH, add the same amounts of acid/hydroxide solution.

NOTE 2 If the pH is less than 5, pH-adjustment with calcium carbonate or sodium hydrogen carbonate (4.6) is preferable. This will also improve the buffer capacity. The amount added should be chosen so that a carbonate residue is left in the sample even after titration.

Add 1 ml of potassium chromate indicator solution (4.2). Titrate the solution by dropwise addition of silver nitrate solution until the colour of the solution just changes to a reddish brown (volume V_s).

After addition of one drop of sodium chloride solution (4.3), the colour should disappear.